

AS91167 – Demonstrate understanding of oxidation-reduction

Keywords

balanced	number of atoms for each element in the reaction and the total charge are the same for both the reactants and the products
disproportionation	redox reaction in which the same element is both oxidized and reduced.
electron transfer	loss or gain of electrons between atoms / ions
half equation	electron-half-equations / half-equations / ionic-half-equations / half-reactions: Electrons are being lost (oxidation) or electrons are being gained (reduction): any redox reaction is made up of two half-reactions
halide	ion formed by halogen eg chloride Cl^- , bromide Br^- , iodide I^-
halogen	non-metal element from group 17 e.g. chlorine Cl_2 , bromine Br_2 , iodine I_2
ion	an atom or a group of atoms that has acquired a charge by gaining or losing one or more electrons e.g. dichromate $\text{Cr}_2\text{O}_7^{2-}$ or Zn^{2+}
observations	changes that can be seen, heard or felt e.g. a colour change or gas bubbles
oxidant	a substance that oxidises another substance; an oxidising agent
oxidation number	oxidation state; sort of electronic book-keeping system to help work out what has been oxidised or reduced
oxidation	loss of electrons; an increase in oxidation number
oxidising agent	a substance that oxidises another substance; an oxidant; substance that removes electrons from other species; it gets reduced in the process
product	substance resulting from a chemical reaction (right side of the \rightarrow)
reactant	substance participating in a chemical reaction (left side of the \rightarrow)
redox reactions	electron transfer reactions are also known as oxidation–reduction reactions
reducing agent	a substance that reduces another substance; a reductant; substance that donates electrons to other species; it gets oxidised in the process
reductant	a substance that reduces another substance; a reducing agent
reduction	gain of electrons; a decrease in oxidation number

