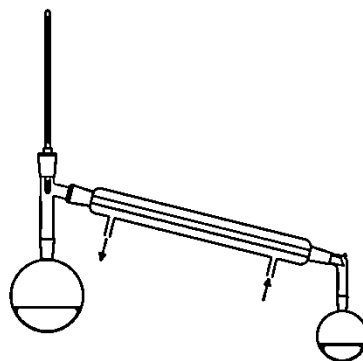
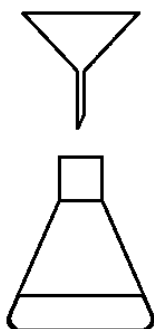


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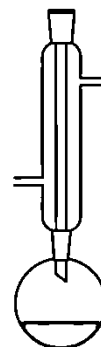


**The separating funnel** allows two reactants to be shaken together and for compounds to be separated based on their solubility or polarity.

E.g. an ester or haloalkane is not soluble in water and will not mix with the aqueous layer. The more dense liquid will be in the lower level, and can be run off by removing the stopper and opening the tap. The less dense layer is poured off through the top of the separating funnel.



**Filtering.** The drying agent e.g. solid anhydrous sodium sulfate absorbs any water left in the sample and can be separated / removed by filtration. E.g. removal of the solid drying agent after drying the organic layer in the preparation of an ester or a haloalkane.



**Refluxing.** This is used to heat an organic reaction (to speed it up or make it take place) without losing volatile organic reactants or products.

**Addition of a carbonate.** (sodium carbonate solution / sodium hydrogen carbonate solution) This is to remove / neutralise any acid e.g. conc.  $\text{H}_2\text{SO}_4$  and unreacted carboxylic acid in ester production, or unreacted conc.  $\text{HCl}$  in the preparation of a haloalkane from a tertiary alcohol.

Extra notes: