



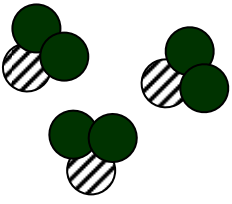
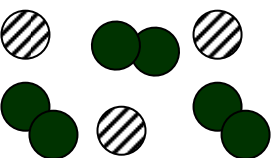


The test to identify CO_2 carbon dioxide gas	The test to identify H_2 hydrogen gas	The test to identify O_2 oxygen gas	Reactants used to prepare hydrogen in the lab
Limewater turns milky	The gas burns with a squeaky pop	The gas will relight a glowing splint	Zinc & hydrochloric acid
Reactants used to prepare oxygen in the lab	Reactants used to prepare carbon dioxide in the lab	Method for collecting a sample of hydrogen gas	Method for collecting a sample of oxygen gas
Hydrogen peroxide & manganese dioxide	Calcium carbonate & hydrochloric acid	Upward delivery OR displacement of water	Displacement of water
Arrange these gases in order of increasing density: CO_2 H_2 O_2	To collect a gas by displacement of water it must be almost ____ in water	Method for collecting a sample of carbon dioxide gas	Name for the starting materials in a chemical reaction
H_2 O_2 CO_2	insoluble	Downward delivery OR displacement of water	reactants
Gas that puts the fizz in the fizzy drink, and is used in fire extinguishers	Lighter than air gas used in weather balloons & the manufacture of margarine and fertilisers	Gas used to breathe in space, under water or in hospitals	Name for the substances made in a chemical reaction
Carbon dioxide	hydrogen	oxygen	products

A more technical word for burning	Bubbling Temperature change Colour change Are signs of:-	Change in appearance New smell Something disappears Are signs of:-	Melting, solidifying, evaporating, condensing Are signs of:-
combustion	A chemical change	A chemical change	A physical change
In a chemical change a *** substance is made	What changes may be easier to reverse, physical or chemical ones?	What kind of equation is this? Magnesium + hydrochloric acid → magnesium chloride + hydrogen	Give the chemical formulae for hydrochloric and sulfuric acid
new	physical	Word equation	HCl H ₂ SO ₄
What are the 3 subatomic particles found in the atom?	When a candle is burnt under a jar, give 2 reasons why it goes out.	What 2 things are needed in addition to iron, for it to rust?	What is the special name given the corrosion of iron (and steel)
Proton, neutron and electron	Runs out of O ₂ The CO ₂ made puts the flame out	Water & oxygen	Rusting
Describe how sodium metal appears when it is freshly sliced	Describe how sodium metal reacts with water	Why does steel wool burn more fiercely in oxygen than in air?	What is the difference between a dilute and a concentrated acid?
Shiny	Floats, fizzes, moves, melts to a ball, may catch fire, yellow flame	Air has only about 21% O ₂ in it	The proportion of acid / water

Name 3 substances a homemade acid-base indicator could be made from	What colour would red and blue litmus be if dipped in acid?	What colour would red and blue litmus be if dipped in alkali?	A student dips a piece of red litmus into a liquid and it stays red. Is the liquid an acid?
Red cabbage, flower petals, tea etc	Red stays red Blue turns red	Red turns blue Blue stays blue	It could be an acid or it could be neutral
Identify this symbol 	Identify this symbol 	Identify this symbol 	Identify this symbol 
Corrosive	(h) Harmful / (i) irritant	Toxic	Highly flammable
Metal + acid → salt + ***	Metal carbonate + acid → salt + *** + ***	What type of pollution can the burning of fossil fuels produce?	What word describes the cancelling out of acidity by an alkali (or vice versa)
Hydrogen	Water & carbon dioxide	Acid rain	Neutralisation
What scale goes from 0-14 and describes how acidic or alkaline a substance is?	Arrange these from most to least acidic: citric acid, water, baking soda, car battery acid	What is the pH of a) a strong acid b) a weak acid	What is the pH of a) a strong alkali b) a weak alkali
pH scale	battery acid / citric / water / baking soda	a) 0-3 b) 4-6	a) 11-14 b) 8-10

A piece of universal indicator was dipped in a liquid and went green. What is the pH?	Arrange these from most to least acidic: vinegar, milk, oven cleaner, orange juice	Give at least 3 observations for when a piece of Mg is dropped into hydrochloric acid.	What 2 properties make CO ₂ suitable for use in fire extinguishers
7 / neutral	Vinegar / o-juice / milk / oven cleaner	Bubbles, Mg disappears, tube gets hot etc	Heavier than air, does not let things burn in it
Why can oxygen gas NOT be collected by upward or downward delivery methods?	Why does 2 mol L ⁻¹ HCl react faster with Mg than 0.5 mol L ⁻¹ HCl?	Only iron (and steel) RUST. What word is used for all other metals?	What substances are formed when a candle burns as well as water?
Its about the same density as air	More acid particles that can collide with the Mg	Corrode / corrosion	carbon dioxide and C (soot) & some carbon monoxide
What word means the breaking down of a chemical eg when hydrogen peroxide turns in water & oxygen	Element, compound or mixture? 	Element, compound or mixture? 	What is a molecule? 
Decomposition	Mixture (of an element & a compound)	Compound	2 or more elements chemically joined together
Element, compound or mixture? 	Element, compound or mixture? 	Element, compound or mixture? 	Element, compound or mixture? 
Element	Compound	Element	Mixture (of 2 elements)