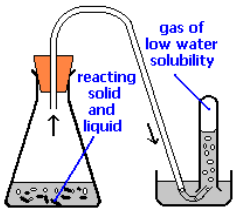
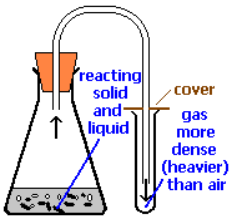
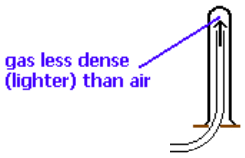
A sour tasting substance which neutralises a base and has a pH of less than 7	Rain which contains dissolved air pollutants eg.CO_2, & NO_2 . It damages plants and buildings	Bases which are soluble in water	Number of protons in the nucleus
Acid	**Acid rain**	**Alkalis**	**Atomic number**
Particles too small to see. Made up of protons, neutrons and electrons	A soapy substance which neutralises an acid and has a pH of more than 7	Positively charged part of the atom, found inside the nucleus	Changes in which one or more new substances are formed.
Atom	**Base**	**Proton**	**Chemical reaction**
A solution which contains large amounts of dissolved solute	Name for the ability to "eat away" a material eg. Some acids are this	Chemical bond formed by the sharing of electrons between two atoms	A solution which contains small amounts of dissolved solute
Concentrated	**Corrosive**	**Covalent bond**	**Dilute**
Pure substances made up of only one type of atom, cannot be broken down in a chemical reaction	Substance which turns different colours in acidic and basic solutions. eg. litmus	Chemical bond resulting from the attraction between oppositely charged ions	Negatively charged particle of an atom, located in energy levels (shells) outside the nucleus.
Element	**Indicator**	**Ionic bond**	**Electrons**

A chemical formula showing molecular composition eg. H_2O	Tiny particles containing two or more atoms in a fixed ratio and joined by chemical bond	Reaction where an acid cancels out a base producing a salt and water	List of elements in order of their atomic number
Molecular formula	**Molecule**	**Neutralisation reaction**	**Periodic table**
Number scale (0-14) indicating how acidic or basic a chemical is. Below 7 = acid, above 7 = base, 7 = neutral	Change where the properties of the substance changes but the substance is still the same	New substances produced in a chemical reaction. Always written on the right hand side of the equation arrow	Substances you start with in a chemical reaction. Always written on the left hand side of the equation arrow
pH scale	**Physical change**	**Product**	**Reactants**
Acids which usually react quickly, breaking up easily into hydrogen ions in solution	Acids which usually react slowly. Do not break up easily into hydrogen ions in solution	**HCl**	**H_2SO_4**
Strong acids	**Weak acids**	**Hydrochloric acid**	**Sulfuric acid**
NaOH	Colourless gas which is 15 times lighter than any other substance. Makes a pop sound when lit	Colourless gas which is needed for respiration. Relights a glowing match	Another word for burning. fuel + oxygen → carbon dioxide + water
Sodium hydroxide	**Hydrogen H_2**	**Oxygen O_2**	**Combustion**

Colourless gas which is heavier than air, needed for photosynthesis. Turns limewater milky	Gas used in fire extinguisher	Gas which is highly explosive and used to make margarine	**Method to collect gas** 
Carbon dioxide CO_2	**Carbon dioxide CO_2**	**Hydrogen H_2**	**Displacement of water**
Number of protons and neutrons in an atom	Colour that blue litmus is in an acid	Colour that red litmus is in a base	Colour that red litmus is in an acid
Mass number	**Red**	**Blue**	**Red**
Colour that blue litmus is in a base	Colour of universal indicator in a strong acid eg. HCl	Colour of universal indicator in a weak acid eg. lemon juice	Colour of universal indicator in a neutral solution eg. water
Blue	**Red**	**Orange**	**Green**
Colour of universal indicator in a weak base eg. tooth paste	Colour of universal indicator in a strong base eg. NaOH	pH of less than 7	pH of 7
Blue	**Purple**	**Acid**	**Neutral**

pH of more than 7	The gas which relights a glowing splint	The gas which is used in weather balloons	Rusting is a form of corrosion where iron reacts with ……
Basic / alkaline	**Oxygen**	**Hydrogen**	**Water and oxygen**
Boiling water removes …… from water so iron wont rust	Metal + oxygen → metal oxide Eg. Aluminium + oxygen →	Most metals react with ……	Indicators can be made from
Oxygen	**Aluminium oxide**	**Acid and air**	**Coloured flowers, red cabbage etc**
Method to collect carbon dioxide 	**Method to collect hydrogen** 	Corrosion is a name when metals react with oxygen to make a ……	Signs of a chemical reaction include …..
Upward displacement of air	**Downward displacement of air**	**Metal oxide**	**Bubbles, colour change, temperature change etc.**
Melting butter, drying cloths, water cycle are all examples of ……	Burning Mg ribbon, cooking, fire works, bleaching cloths are all examples of ……	Chemical formula of lime water or calcium hydroxide is ……	Particle of the atom which has no charge, located in the nucleus of the atom
Physical change	**Chemical change**	**$Ca(OH)_2$**	**Neutrons**
