


Material that prevents radioactive emission from passing through it	Number of radioactive emissions/counts per second - also known as "rate of decay"	Release of alpha particle from unstable nucleus (a 2 ⁺ helium ion or a helium nucleus)	Number of protons in a nucleus
Absorber	Activity	Alpha emission	Atomic number
Natural radiation from everyday surroundings	The way charged particles are forced to move in a magnetic field	Release of beta particle from unstable nucleus (a single electron)	The number of counts on a Geiger counter (each count is caused by radiation causing ionisation)
Background radiation	Behaviour in a magnetic field	Beta emission	Count rate
Atoms are the smallest building blocks of matter and are indivisible (can't be split up)	Particles close together	Negatively charged subatomic particle	Release of gamma wave from unstable nucleus (Electromagnetic spectrum)
Dalton's model of the atom	Dense	Electron	Gamma emission
Device to measure radioactivity	Rutherford experiment where he bombarded a thin gold foil with alpha particles in a vacuum	Time taken for half the radioactive nuclei in a sample to decay	Remove electrons from orbit by collision
Geiger counter	Gold foil experiment	Half life	Ionisation

A measure of the ability to remove electrons	Same number of protons, different number of neutrons or same element but different number of neutrons	A sheet of lead used to absorb alpha and beta and limit gamma	Number of protons and neutrons
Ionising ability	Isotope	Lead screen	Mass number
Charge that attracts positive charges and repels negative charges	Uncharged subatomic particle	An equation to show a radioactive change with balanced atomic and mass numbers	Small, dense centre of atom containing neutrons and protons
Negative	Neutron	Nuclear equation	Nucleus
Ability of radioactive emissions to pass through materials	A neutral atom of positive charge with negative charges embedded into it	Charge that repels positive charges and attracts negative charges	Positively charged subatomic particle
Penetration ability	Plum pudding model	Positive	Proton
Emission of α , β or γ in order to transmute (change contents of nucleus) or to lose excess energy	Source (point of origin) of radioactive emissions	Number of radioactive emissions/counts per second - also known as "activity"	Neutrons and protons in a small dense nucleus with orbiting electrons and large gaps within the atom
Radioactive decay	Radioactive source	Rate of decay	Rutherford's model of the atom

<p>The way the model suggests the atom is arranged (Dalton, Thomson, Rutherford)</p>	<p>The plum pudding model</p> 	<p>Alpha, beta or gamma</p>	<p>Nucleus that will emit radiation</p>
<p>Structure of the atom</p>	<p>Thomson's model of the atom</p>	<p>Type of radiation</p>	<p>Unstable nuclei</p>
<p>Chamber with no particles inside it</p>			
<p>Vacuum chamber</p>			