


91167 REXOX FLASH CARDS

Cut out and fold to produce cards, as indicated below.

Note: there are 13 oxidants & 11 reductants

<p>OXIDANT</p> <p>Purple permanganate MnO_4^- (acidified)</p> <p>is reduced to</p>	 <p>FOLD HERE & GLUE BACK TO BACK</p>
<p>the manganese(II) ion</p> <p>Mn^{2+}</p> <p>which is colourless</p>	

<p>OXIDANT</p> <p>colourless $O_2(g)$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>orange-brown $I_2(aq)$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>yellow green chlorine gas $Cl_2(g)$</p> <p>OR</p> <p>chlorine water $Cl_2(aq)$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>colourless $H^+(aq)$ (acid)</p> <p>is reduced to</p>
<p>the oxide ion</p> <p>O^{2-}</p> <p>colour depends...</p> <p>MgO white</p> <p>CuO black</p>	<p>the iodide ion</p> <p>$I^-(aq)$</p> <p>which is colourless</p>	<p>the chloride ion</p> <p>$Cl^-(aq)$</p> <p>which is colourless</p>	<p>hydrogen gas</p> <p>$H_2(g)$</p> <p>which is colourless</p>
<p>OXIDANT</p> <p>pale orange iron(III) $Fe^{3+}(aq)$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>colourless hydrogen peroxide $H_2O_2(aq)$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>purple permanganate $MnO_4^-(aq)$ (acidified)</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>orange dichromate $Cr_2O_7^{2-}(aq)$ (acidified)</p> <p>is reduced to</p>
<p>the iron(II) ion</p> <p>$Fe^{2+}(aq)$</p> <p>which is pale green</p>	<p>water</p> <p>$H_2O(l)$</p> <p>which is colourless</p>	<p>the manganese(II) ion</p> <p>$Mn^{2+}(aq)$</p> <p>which is colourless</p>	<p>the chromium(III) ion</p> <p>$Cr^{3+}(aq)$</p> <p>which is green</p>

<p>REDUCTANT</p> <p>metals eg pinky orange Cu(s) & grey Mg(s) and Fe(s)</p> <p>are oxidised to</p>	<p>REDUCTANT</p> <p>black C(s)</p> <p>is oxidised to</p>	<p>REDUCTANT</p> <p>colourless CO(g)</p> <p>is oxidised to</p>	<p>REDUCTANT</p> <p>colourless hydrogen gas H₂(g)</p> <p>is oxidised to</p>
<p>metal ions(aq)</p> <p>colour depends...</p> <p>Cu²⁺ blue</p> <p>Mg²⁺ colourless</p> <p>Fe²⁺ pale green</p> <p>Fe³⁺ pale orange</p>	<p>either carbon monoxide or carbon dioxide gas</p> <p>CO(g) or CO₂(g)</p> <p>which are colourless</p>	<p>carbon dioxide gas</p> <p>CO₂(g)</p> <p>which is colourless</p>	<p>H⁺(aq)</p> <p>which is colourless</p>
<p>REDUCTANT</p> <p>pale green iron(II) ion Fe²⁺(aq)</p> <p>is oxidised to</p>	<p>REDUCTANT</p> <p>colourless bromide ion, Br⁻(aq)</p> <p>is oxidised to</p>	<p>REDUCTANT</p> <p>colourless iodide ion, I⁻(aq)</p> <p>is oxidised to</p>	<p>REDUCTANT</p> <p>colourless SO₂(g) gas</p> <p>is oxidised to</p>
<p>the iron (III) ion Fe³⁺</p> <p>which is a pale orange</p>	<p>bromine solution Br₂</p> <p>which is a red-brown colour</p>	<p>iodine solution I₂(aq)</p> <p>which is an orange-brown colour</p>	<p>the sulfate ion SO₄²⁻(aq)</p> <p>which is colourless</p>

<p>REDUCTANT</p> <p>colourless hydrogen sulfite $\text{HSO}_3^-(\text{aq})$ or sulfite $\text{SO}_3^{2-}(\text{aq})$</p> <p>is oxidised to</p>	<p>REDUCTANT</p> <p>colourless hydrogen peroxide $\text{H}_2\text{O}_2(\text{aq})$</p> <p>is oxidised to</p>	<p>REDUCTANT</p> <p>colourless hydrogen sulfide gas $\text{H}_2\text{S}(\text{g})$</p> <p>is oxidised to</p>	<p>OXIDANT</p> <p>colourless iodate ion, $\text{IO}_3^-(\text{aq})$</p> <p>is oxidised to</p>
<p>the sulfate ion $\text{SO}_4^{2-}(\text{aq})$</p> <p>which is colourless</p>	<p>oxygen gas $\text{O}_2(\text{g})$</p> <p>which is colourless</p>	<p>sulfur $\text{S}(\text{s})$</p> <p>which is yellow</p>	<p>iodine solution $\text{I}_2(\text{aq})$</p> <p>which is an orange-brown colour</p>
<p>OXIDANT</p> <p>red-brown "bromine water" $\text{Br}_2(\text{aq})$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>colourless hydrogen peroxide $\text{H}_2\text{O}_2(\text{aq})$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>colourless chlorite $\text{OCl}^-(\text{aq})$</p> <p>is reduced to</p>	<p>OXIDANT</p> <p>colourless conc. nitric acid HNO_3 (may have a brown tinge)</p> <p>is reduced to</p>
<p>the bromide ion $\text{Br}^-(\text{aq})$</p> <p>which is colourless</p>	<p>water $\text{H}_2\text{O}(\text{l})$</p> <p>which is colourless</p>	<p>the chloride ion $\text{Cl}^-(\text{aq})$</p> <p>which is colourless</p>	<p>nitrogen dioxide gas NO_2</p> <p>which is brown</p>

