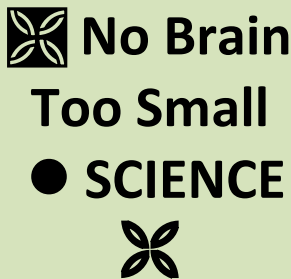


metal + oxygen →	<u>Example</u> magnesium + oxygen →	metal + water →	<u>Example</u> sodium + water →
metal oxide	magnesium oxide	metal hydroxide + hydrogen	sodium hydroxide + hydrogen
metal + acid →	<u>Example</u> zinc + sulfuric acid →	<u>Example</u> magnesium + hydrochloric acid →	acid + base →
salt + hydrogen	zinc sulfate + hydrogen	magnesium chloride + hydrogen	salt + water
<u>Example</u> sulfuric acid + sodium hydroxide →	<u>Example</u> hydrochloric acid + copper oxide →	a base is a...	a metal oxide or metal hydroxide is called ...
sodium sulfate + water	copper chloride + water	metal oxide or metal hydroxide	a base
sulfuric acid has the formula	hydrochloric acid has the formula	sulfuric acid forms salts called...	hydrochloric acid forms salts called...
H_2SO_4	HCl	sulfates	chlorides

acid + carbonate →	acid + hydrogen carbonate →	the visible sign of reaction between metal + acid is....	the visible sign of reaction between a carbonate + acid is....
salt + water + carbon dioxide	salt + water + carbon dioxide	Fizzing (due to H ₂)	Fizzing (due to CO ₂)
<u>Example</u> sulfuric acid + copper carbonate →	<u>Example</u> hydrochloric acid + zinc carbonate →	<u>Example</u> sulfuric acid + sodium hydrogen carbonate →	the SO ₄ ²⁻ ion is called the ___ ion
copper sulfate + water + carbon dioxide	zinc chloride + water + carbon dioxide	sodium sulfate + water + carbon dioxide	sulfate
the CO ₃ ²⁻ ion is called the ___ ion	the HCO ₃ ⁻ ion is called the ___ ion	the OH ⁻ ion is called the ___ ion	the NO ₃ ⁻ ion is called the ___ ion
carbonate	hydrogen carbonate (also bicarbonate)	hydroxide	nitrate
a lit splint is held at the mouth of the test tube - the gas that ignites with a "pop" is...	the gas is bubbled through colourless limewater which turns cloudy / milky - the gas is...	a base that is soluble in water is called an ___	pH < 7 is ___, pH 7 = ___, & pH > 7 = ___
hydrogen, H ₂	carbon dioxide, CO ₂	alkali	acidic neutral alkaline/basic

nitric acid has the formula....	nitric acid forms salts called ...	<u>Example</u> nitric acid + zinc oxide →	<u>Example</u> nitric acid + copper carbonate →
HNO_3	nitrates	zinc nitrate + water	zinc nitrate + water + carbon dioxide
which ion do all acids contain?	which ion do all alkalis contain?	what is the name for a base that is soluble in water?	what is the formula of the hydroxide ion?
H^+ (or H_3O^+)	OH^-	alkali	OH^-
what is the pH of a neutral solution?	what pH are soluble bases?	what colour is universal indicator with a strong acid?	what colour is universal indicator with a weak acid?
7	>7	red / pink	orange / yellow
what colour is universal indicator with a neutral solution?	what colour is universal indicator with a weak alkali?	what colour is universal indicator with a strong alkali?	what is the pH of a weak acid?
green	blue	purple	3 - 6

what is the pH of a strong acid?	what is the pH of a weak alkali?	what is the pH of a strong alkali?	what colour change is seen if blue litmus added to acid?
0 - 2	8 - 11	12 - 14	blue to red
acid + hydroxide →	<u>Example</u> nitric acid + sodium hydroxide →	<u>Example</u> copper hydroxide + sulfuric acid →	what colour change is seen if blue litmus added to alkali?
salt + water	sodium nitrate + water	copper sulfate + water	no change / remains blue
pH of a liquid that has no effect on blue or red litmus paper is...	acids & alkalis react with each other, the alkali cancels out the acid. This is called ...	combination of hydrogen ions with hydroxide ions to form water is called...	colour of copper(II) oxide
7	neutralisation	neutralisation	black
colour of copper(II) carbonate	colour of copper(II) sulfate	formed when an acid is neutralised by a base or carbonate	
green	blue	a salt	